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Modern Chemistry 1 Stoichiometry Name: \_\_\_\_\_ Class: \_\_\_\_\_ Date: \_\_\_\_\_ CHAPTER 9 REVIEW Stoichiometry SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. \_\_\_\_\_ The coefficients in a chemical equation represent the (a) masses in grams of all reactants and products.

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### Review Stoichiometry Section 1 And 2 Answers

Chapter 9 Review Stoichiometry Section CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N<sub>2</sub> are mixed with 12.0 mol of H<sub>2</sub> according ...

### Chapter 9 Review Stoichiometry Section 1 Answers

Review Vocabulary reactant: the starting substance in a chemical reaction New Vocabulary stoichiometry mole ratio SECTION 1 Defining Stoichiometry 368 Chapter 11 • Stoichiometry Program: Chemistry Component: SE PDF Vendor: Symmetry National Chapter 11 Charles D. Winters/Photo Researchers 0368\_0372\_C11\_S1\_896405.indd 368 2/10/11 11:24 AM

### CHAPTER 11 Stoichiometry

Stoichiometry. SECTION 1. SHORT ANSWER Answer the following questions in the space provided. 1. \_\_\_\_\_ The coefficients in a chemical equation represent the (a) masses in grams of all reactants and products. (b) relative number of moles of reactants and products. (c) number of atoms of each element in each compound in a reaction.

### CHAPTER 9 REVIEW

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Stoichiometry. SECTION 2. PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. The following equation represents a laboratory preparation for oxygen gas: 2KClO<sub>3</sub>(s) → 2KCl(s) + 3O<sub>2</sub>(g) How many moles of O<sub>2</sub> form if 3.0 mol of KClO<sub>3</sub> are totally consumed? 2. Given the following equation: H<sub>2</sub>(g) + F<sub>2</sub>(g) → 2HF(g)

### CHAPTER 9 REVIEW

CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N<sub>2</sub> are mixed with 12.0 mol of H<sub>2</sub> according to the following equation: N

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1 mol Al<sub>2</sub>O<sub>3</sub> or 1 mol Al<sub>2</sub>O<sub>3</sub> \_\_\_ 101.96 g Al<sub>2</sub>O<sub>3</sub> 26.98 g Al \_\_\_ 1 mol Al or \_\_\_ 1 mol Al 26.98 g Al 32.00 g O<sub>2</sub> \_\_\_ 1 mol O<sub>2</sub> or 1 mol O<sub>2</sub> \_\_\_ 32.00 g O<sub>2</sub> To find the number of grams of aluminum equivalent to 26.0 mol of aluminum, the calculation would be as follows. 26.0 mol Al × 26.98 g Al \_\_\_ 1 mol Al = 701 g Al Stoichiometry 291 SECTION 1 ...

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### Modern Chemistry Chapter 9 Homework 9 1 Answers

1. Write the definition of reaction stoichiometry in your own words. Introduction to Stoichiometry SECTION 9.1 amount of given substance (mol) convert into amount of unknown substance (mol) Ratios of substances in chemical reactions can be used as conversion factors. Reaction stoichiometry problems can be approached by looking

### SECTION 9.1 Introduction to Stoichiometry

SECTION 2 continued Date Class \_\_\_\_\_ 60.2 9 42.1 1 a. \ tt mash 01 ox aen Cas i pridui.ed it 100. of lithium c a C ti. I o c. i o g di I ClO c — L Ci(,; — h. The oxygen gas produced in part a has density of 1.43 g/L calculate the volume of this gas. 76 STOICHIOMETRY MODERN CHEMISTRY a. —. 81 g 6. A car

air bag requires 70. L of nitrogen gas ...

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