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Properties Of **Review**

Chapter 16:

Solutions 16.1

Properties of Solutions Solution

Formation The compositions of the solvent and the solute determine whether a substance will dissolve.

The factors that determine how fast a substance dissolves are stirring (agitation) temperature the surface area of the dissolving particles 16.1 Solution

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Formation Stirring and
Solution

Formation Stirring
speeds up the
dissolving process
because fresh solvent
(the water in tea) is
continually brought
into contact with the
surface of the ...

Chapter 16: **Solutions 16.1** **Properties of** **Solutions - [PPT ...**

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- saturated solution
- solubility -unsaturated solution -miscible
- immiscible
- supersaturated solution -henry's law

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Solutions are likely to have properties similar to those of their major component—usually the solvent. However,

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some solution properties differ significantly from those of the solvent. Here, we will focus on liquid solutions that have a solid solute, but many of the effects we will discuss in this section are applicable to all solutions.

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problem. 161 ...

Section 161 **Properties Of** **Solutions Pages 471** **477**

Chapter 16 Solutions
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167 SECTION 16.1
PROPERTIES OF
SOLUTIONS (pages
471-477) This section
identifies the factors
that affect the
solubility of a
substance and
determine the rate at
which a solute
dissolves. Solution
Formation (pages
471-472) Look at
Figure 16.1 on page
471 to help you answer
Questions 1 and 2. 1.

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Properties Of **05 Chem GRSW Ch16.SE/TE**

Properties of Solution.
Different properties of solutions are as follows: It is a homogeneous mixture. Its particles are too tiny and have a diameter less than 1 nm. The particles are not visible to naked eyes. Particles don't scatter a beam of light passing through it and hence the path of the light is not visible.

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Solutions - Definition, Properties, Types, Videos & Examples

This section identifies the factors that affect the solubility of a substance and determine the rate at which a solute dissolves. Solution Formation (pages 471-472) Look at Figure 16.1 on page 471 to help you answer Questions 1 and 2. 1.

Underline

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Section 161
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**(PDF) SECTION 16.1
PROPERTIES OF
SOLUTIONS (pages
471-477 ...**

SECTION 16.1

PROPERTIES OF

SOLUTIONS 1. The

solubility of CO_2 in
water at 1.22 atm is
0.54 g/L. What is the

solubility of carbon
dioxide at 1.86 atm?

Assume that

temperature is

constant. 2. What mass
of KCl will produce a

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Solutions

saturated solution in
500.0 g of water at 20
C?

SECTION 16.1 PROPERTIES OF SOLUTIONS

24) A solution containing 10.0 g of an unknown liquid and 90.0 g water has a freezing point of -3.33°C . Given $K_f = 1.86^{\circ}\text{C}/m$ for water, the molar mass of the unknown liquid is _____ g/mol. A)161 B)69.0

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Solutions
C)62.1 D)333 E)619

25) A solution is prepared by dissolving 0.60 g of nicotine (a nonelectrolyte) in water to make 12 mL of solution.

A.P. Chemistry Practice Test: Ch. 11, Solutions MULTIPLE ...

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Chapter 16

1977—Pub. L. 95-30 substituted “section 63” for “section 63(a)”.
Effective Date of 1977 Amendment
Amendment by Pub. L. 95-30 applicable to taxable years beginning after Dec. 31, 1976 , see section 106(a) of Pub. L. 95-30 , set out as a note under section 1 of this title .

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Properties Of Solutions **26 U.S. Code § 161 - Allowance of deductions | U.S. Code ...**

The amount of a substance that dissolves in a given quantity of solvent at a given temperature. 2. The solubility of a gas in a liquid is directly proportional to the pressure of the gas above the liquid. 3. Solution that contains the maximum amount of solute for a given

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amount of solvent at a
constant temperature.

4.

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